Factors Affecting Reaction Rates:

Temperature

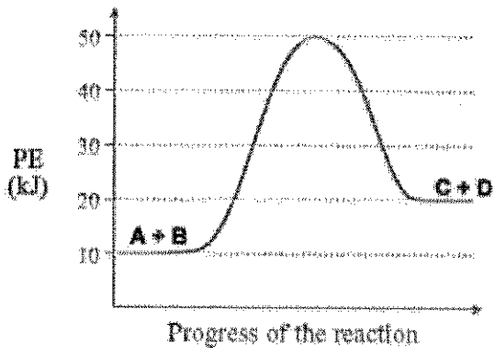
Concentration:

Particle size/surface area

Catalysts:

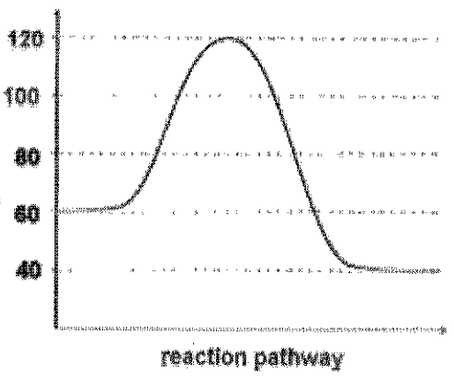
Potential Energy Diagrams

1. Is this reaction endothermic or exothermic? _____



a. What is the energy of the reactants? _____
b. What is the energy of the products? _____
c. What is the activation energy? _____
d. What is ΔH?_____
e. Is energy absorbed or released making the products? _____
f. Indicate with a ✳ the location of the activated complex.

2. Is this reaction endothermic or exothermic? _____



a. What is the energy of the reactants? _____
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d. What is ΔH?_____
e. Is energy absorbed or released making the products? _____
f. Indicate with a ✳ the location of the activated complex.

Chemistry 122- EQUILIBRIUM / REACTION RATES

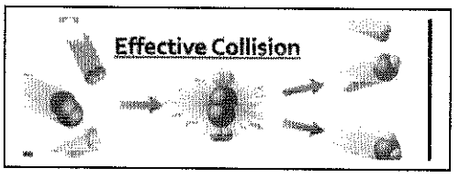
Key terms ☺

COLLISION THEORY

Reaction rate:

Collision Theory:

Only some collisions = reactions...why?



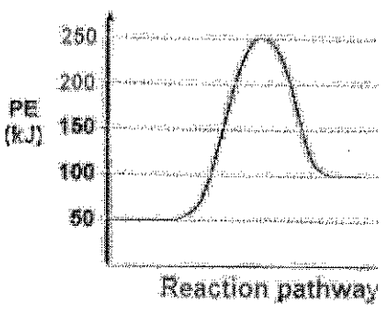
Activation Energy-

Activated complex/transition state-

POTENTIAL ENERGY DIAGRAMS:

Identify reactants, products, activation energy, activated complex and ΔH for each. (pg 543)

Endothermic reactions



Exothermic reactions