

Lewis Dot, Formula Unit & Naming Practice Sheet

Notes:

1. An **ionic bond** is an attraction of a *cation* for an *anion* resulting from the transfer of electrons. Remember, the smaller nonmetals are more electronegative and pull the electrons close, away from the larger, less electronegative metals.
2. When naming ionic compounds, the Metal is named first, followed by the nonmetal with an *-ide* ending. *Ex. Sodium Fluorine becomes Sodium Fluoride.*
3. **Formula Unit:** Lowest whole number ratio of elements in the compound. *Ex. Ca₃N₂*

<p>1. Draw the Lewis Structure for Mg & Cl</p> <p>Formula Unit: _____</p> <p>Name of Compound: _____</p>	<p>2. Draw the Lewis Structure for Mg & S</p> <p>Formula Unit: _____</p> <p>Name of Compound: _____</p>
<p>3. Draw the Lewis Structure for K & F</p> <p>Formula Unit: _____</p> <p>Name of Compound: _____</p>	<p>4. Draw the Lewis Structure for K & O</p> <p>Formula Unit: _____</p> <p>Name of Compound: _____</p>
<p>5. Draw the Lewis Structure for Be & N</p> <p>Formula Unit: _____</p> <p>Name of Compound: _____</p>	<p>6. Draw the Lewis Structure for Ca & P</p> <p>Formula Unit: _____</p> <p>Name of Compound: _____</p>

<p>7. Draw the Lewis Structure for Al & F</p> <p>Formula Unit: _____</p> <p>Name of Compound:</p>	<p>8. Draw the Lewis Structure for Ca & I</p> <p>Formula Unit: _____</p> <p>Name of Compound:</p>
<p>9. Draw the Lewis Structure for Rb & O</p> <p>Formula Unit: _____</p> <p>Name of Compound:</p>	<p>10. Draw the Lewis Structure for Sr & F</p> <p>Formula Unit: _____</p> <p>Name of Compound:</p>
<p>11. Draw the Lewis Structure for Al & Cl</p> <p>Formula Unit: _____</p> <p>Name of Compound:</p>	<p>12. Draw the Lewis Structure for Mg & P</p> <p>Formula Unit: _____</p> <p>Name of Compound:</p>
<p>13. Draw the Lewis Structure for B & O</p> <p>Formula Unit: _____</p> <p>Name of Compound:</p>	<p>14. Draw the Lewis Structure for Be & S</p> <p>Formula Unit: _____</p> <p>Name of Compound:</p>